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Chapter 6
Ionic Bonding
Exercises

Charge on the ion = $12 - 10 = +2$
The ion is magnesium ion, Mg^{2+} . (Mg atom = (2, 8, 2), Mg ion = (2, 8))

Test Yourself 6.1 (page 92)

1.

Ion	Number of protons	Number of neutrons	Number of electrons
Cl^-	17	18	18
Br^+	35	45	34
F^-	9	10	10

2. Charge is -3, ion is N^{3-} .

3. Copper(I) oxide: Cu_2O , Copper(II) sulphate: $CuSO_4$, SO_4^{2-}

Test Yourself 6.2 (page 94)

1. Sodium and bromine, $NaBr$
Iron and bromine, $FeBr_2$ or $FeBr_3$
Sodium and sulphur, Na_2S
Iron and sulphur FeS or Fe_2S_3

2. Sodium oxide: Na_2O
Aluminium nitrate: $Al(NO_3)_3$
Calcium hydroxide: $Ca(OH)_2$
Potassium sulphate: K_2SO_4
Calcium sulphate: $CaSO_4$

Test Yourself 6.3 (page 95)

1. (a) Strontium chloride is an ionic compound. The ionic bonding between Sr^{2+} and Cl^- is very strong and a lot of energy is needed to break the bond.
(b) It is insoluble in organic solvents and conducts electricity when dissolved in water.

2. Magnesium oxide is an ionic compound. It has a high melting point and thus can withstand high temperatures inside the furnace.

3. Each positive ion M^{2+} is placed next to a negative ion X^- . The ratio of M^{2+} ions to X^- ions is 1:2. Hence, the chemical formula of the ionic compound is MX_2 .

Exercise 6 (page 100)

Foundation

1. A 2. B 3. D 4. D

5. H_2 , Cl_2 , H and BH_3^+ (any three)

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